

Revision

Q1 Potassium (at No. 19) and Chlorine (at No. 17) form a compound. Explain the formation of the compound on the basis of a) oxidation b) reduction c) oxidising agent (d) Reducing agent



a) Oxidation:- Oxidation is process in which atom or ion loses an electron
 $K \longrightarrow K^+ + e^-$

b) Reduction:- It is process in which atom or ion accept electron
 $Cl_2 + 2e^- \longrightarrow 2Cl^-$

c) Oxidising agent:- It oxidises other substances either by accepting electron or by providing oxygen or an electro-ve ion.



d) Reducing agent:- It reduces other substances either by providing electron or by providing hydrogen or electrove ion.



Q:-2 Show number of bonds in:-

a) ethene molecule b) ethyne molecule

a) $\begin{array}{c} \text{H} \quad \text{H} \\ | \quad | \\ \text{H}-\text{C}=\text{C}-\text{H} \end{array}$ has one double bond and four single covalent bond.

b) $\text{H}-\text{C}\equiv\text{C}-\text{H}$ has two double bond and two single covalent bond.

Q:- Metal M forms a chloride with the formula MCl_2 . What type of bond in MCl_2 . Write the formula of the compound when M combines with sulphur, oxygen and nitrogen.

Ans:- The bond formed between metal and non-metal is ionic bond.

Bond formed between metal M and Chlorine is ionic bond.

When metal M combines with sulphur M_2S

When metal M combines with oxygen M_2O

When metal M combines with nitrogen M_3N_2

Q:- An element M burns in oxygen to form an ionic comp. MO . Write the formula of the comp. formed if this element is made to combine with chlorine and sulphur separately.

Ans:- MCl_2
 MS