

CLASS- IX
CHAPTER- 2

SUBJECT- CHEMISTRY
TEACHER- ANAMIKA

Revision - CH-2 (Chemical Bonding)

Q1:- Divide the following redox reactions into oxidation and reduction half reaction.

- (i) $\text{Zn} + \text{Pb}^{2+} \rightarrow \text{Zn}^{2+} + \text{Pb}$
 (ii) $\text{Cl}_2 + 2\text{Br}^- \rightarrow \text{Br}_2 + 2\text{Cl}^-$
 (iii) $2\text{Cu}^+ \rightarrow \text{Cu} + \text{Cu}^{2+}$

Ans1:- (i) $\text{Zn} - 2\text{e}^- \rightarrow \text{Zn}^{2+}$ [O]
 $\text{Pb}^{2+} + 2\text{e}^- \rightarrow \text{Pb}$ [R]
 (ii) $2\text{Br}^- - 2\text{e}^- \rightarrow \text{Br}_2$ [O]
 $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$ [R]
 (iii) $\text{Cu}^+ - \text{e}^- \rightarrow \text{Cu}^{2+}$ [O]
 $\text{Cu}^+ + \text{e}^- \rightarrow \text{Cu}$ [R]

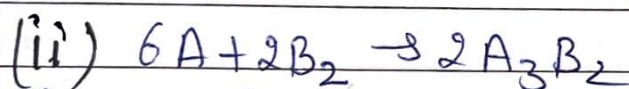
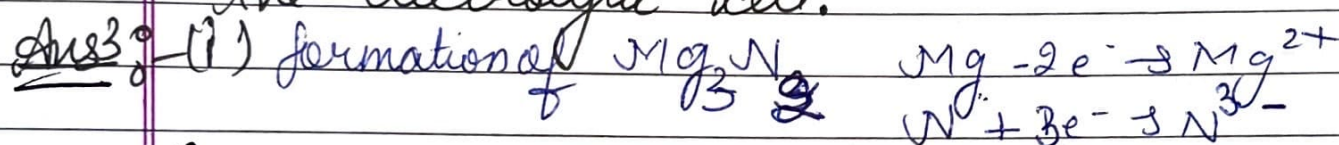
Q2:- Elements A, B & C have atomic number 17, 19 & 10 respectively

- (a) State which one is (i) a non-metal
 (ii) a metal (iii) chemically inert
 (b) Write down the formula of the compound formed by two of the above elements

Ans2:- (a) A = Non-metal
 B = Metal
 C = Chemically inert element

(b) BA

- Q3 Element A has 2 electrons in its M shell. Element B has atomic number 7.
- (a) Write equations to show how A & B form ions.
- (b) If B is a diatomic gas, write the equation for the direct combination of A & B to form a compound.
- (c) If the compound formed by A & B is melted & an electric current is passed through the molten compound, the element A will be obtained at the _____ & B at the _____ of the electrolytic cell.



(iii) cathode, anode

Q4 Give two examples in each case:-

- (i) Solid covalent compounds
(ii) Gaseous non polar compounds
(iii) Liquid non polar compounds.

- Ans 4: (i) Diamond, Silicon carbide
(ii) Methane (CH_4), Ethylene (C_2H_4)
(iii) Gasoline, Toluene.

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