

TENDER HEART HIGH SCHOOL, SEC-33B, CHD

CLASS IX

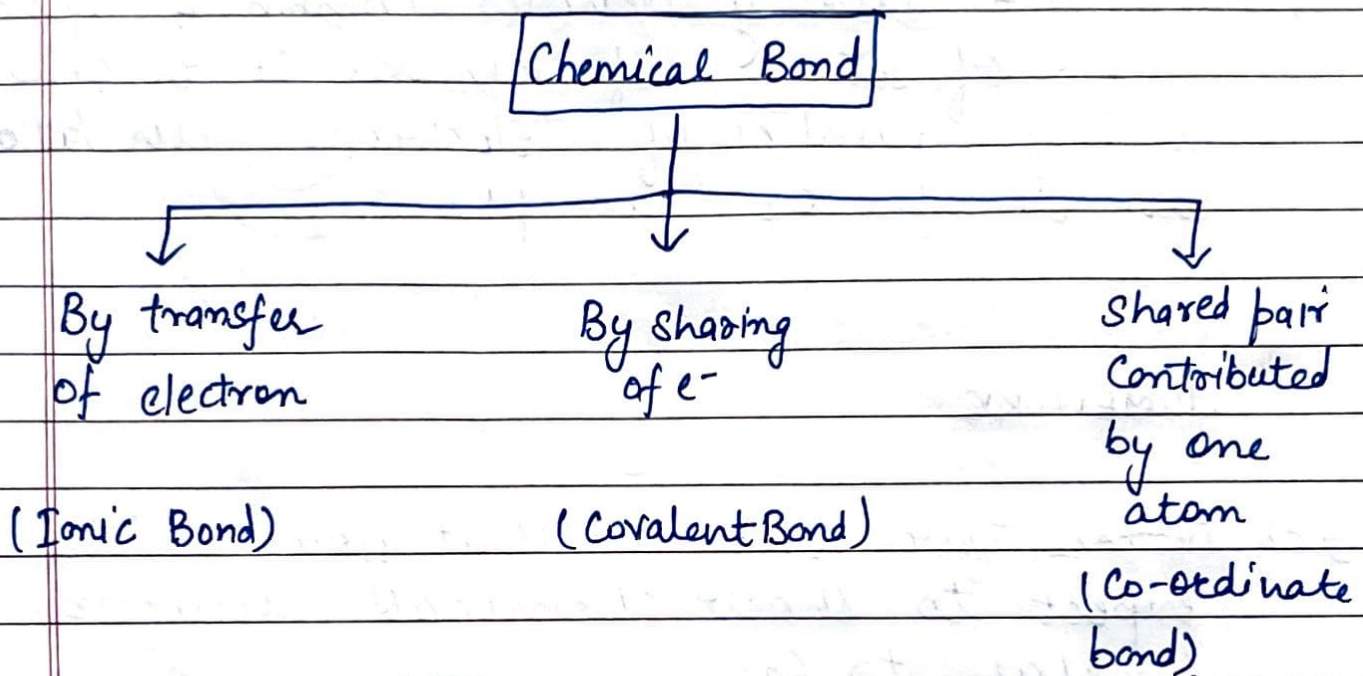
CHAPTER 2

SUBJECT- CHEMISTRY

TEACHER- ANAMIKA

21.10.24

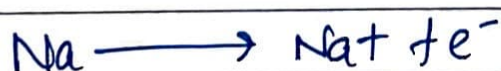
A chemical bond may be defined as force of attraction b/w any two atoms in a molecule to maintain stability

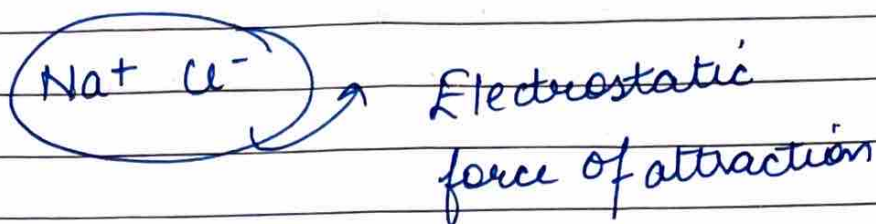


Ionic Bond:- The chemical compound formed as a result of transfer of e^- from one atom of an element to one atom of another element are called ionic compound / Electrovalent bond.

For e.g.

In NaCl



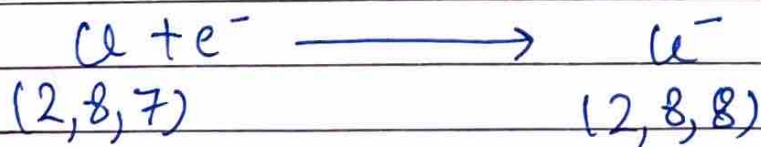
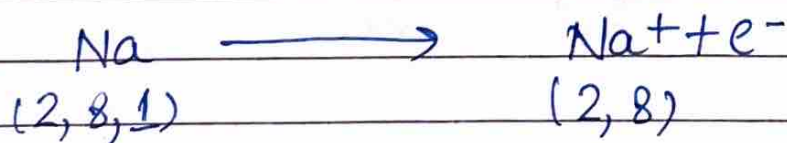


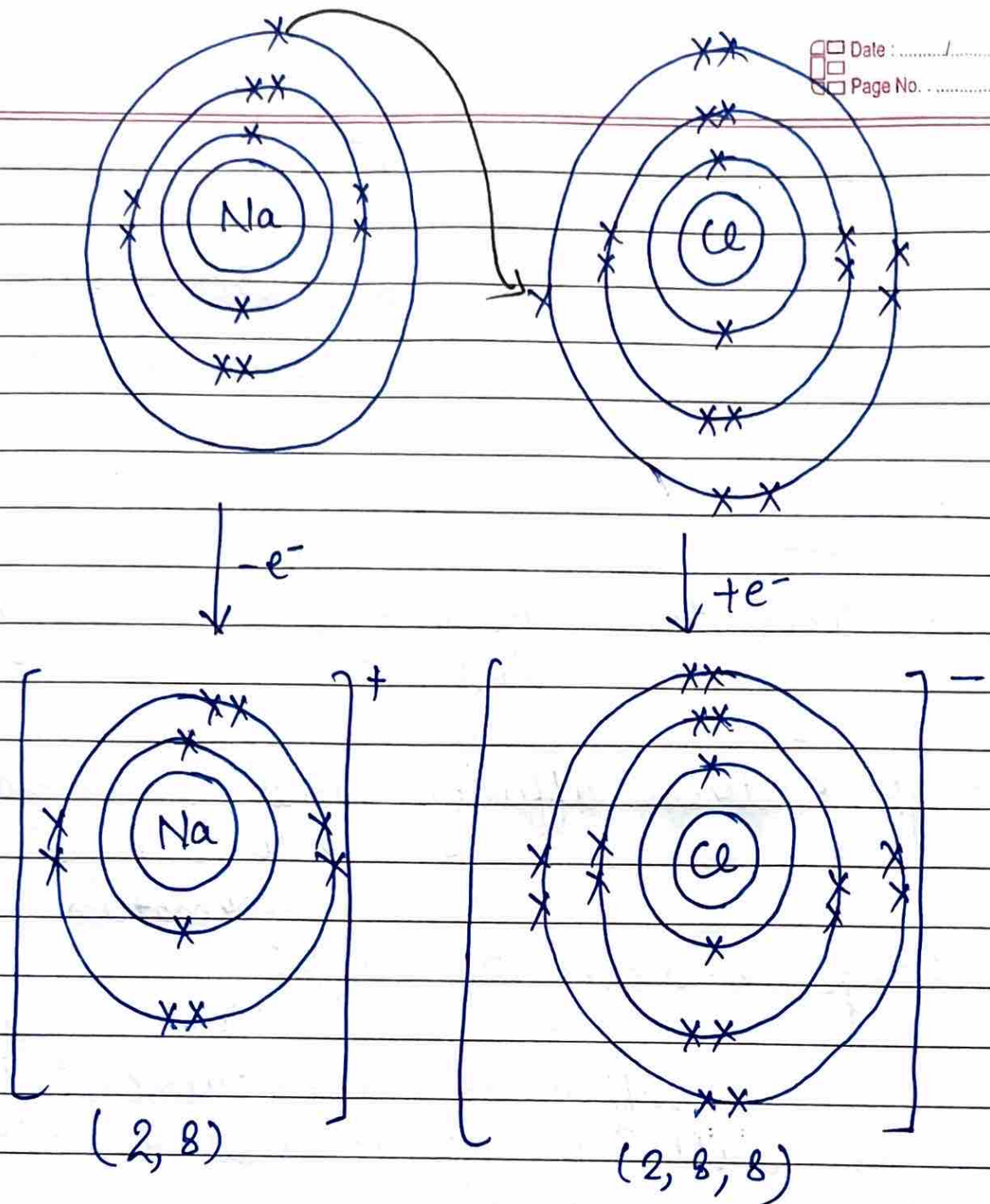
Conditions

1. Low Ionisation potential :- It will easily lose e^-
2. High Electron affinity higher value favours ionic bond formation
3. Large electronegativity diff-
More is the difference in electronegativity more will be ionic nature of resulting compound.

NOTE:- Group 1 Element are more electro +ve
Group 17 " " " electro -ve

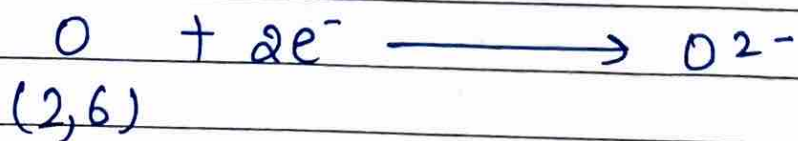
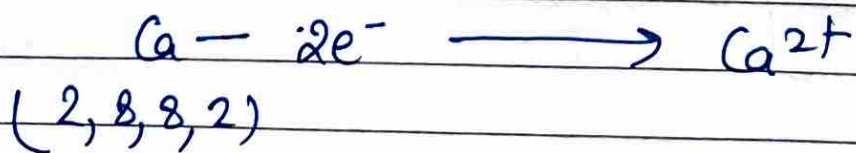
Structure of NaCl:-

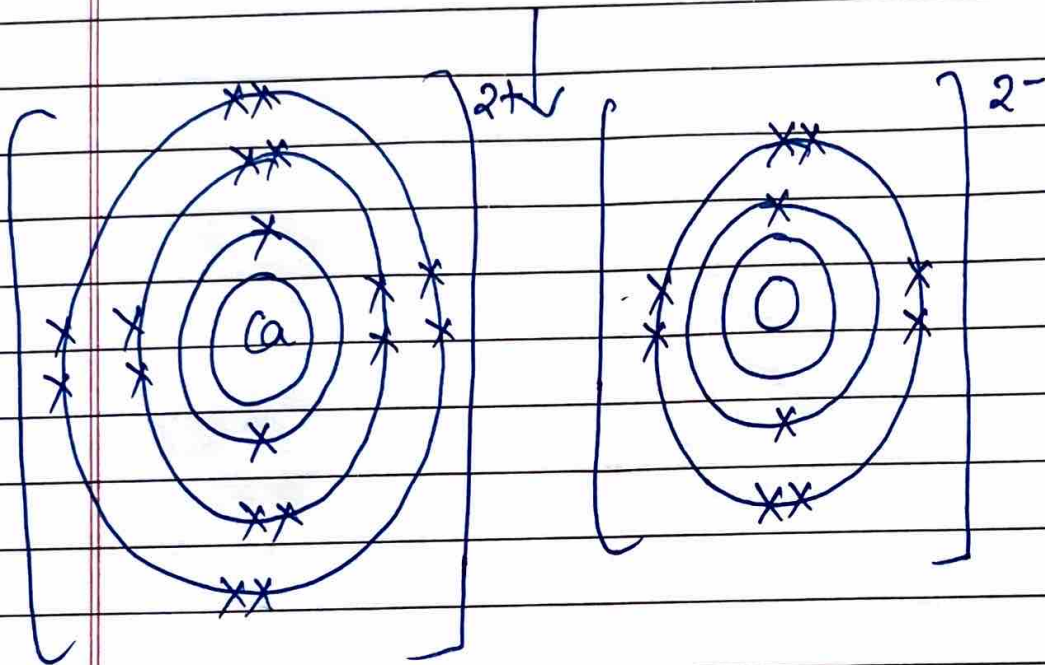
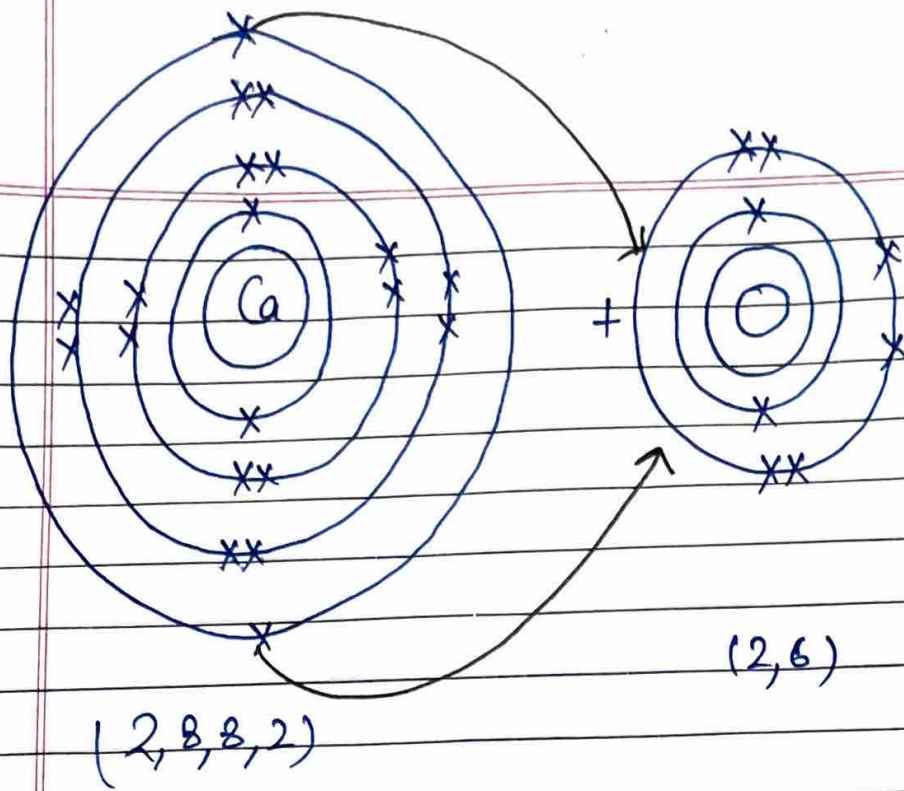




2.

CaO.





NOTE:- In the formation of ionic bond, transfer of e^- involved.

Electro+ve atom undergoes oxidation,
electro-ve atom undergoes reduction.
This is known as redox Reaction.