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TENDER HEART HIGH SCHOOL, SEC-33B
CHD.

Class - VIII

Chapter - 6

Subject - Chemistry

Teacher - Anamika

Exercise 6(A).

Q:-1 Justify position of hydrogen in the periodic table?

Ans:- It belongs to first group and first period of periodic table.

Q:-2 Why does hydrogen show dual nature?

Ans:- Hydrogen shows dual nature since it resembles the alkali metals of group IA and halogens of group VII A (17)

Q:-3 Compare hydrogen with alkali metals on the basis of-

(1) Ion formation:- Each of them form cation, i.e. positive ion by loss of e^- .

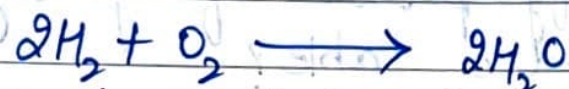


2. Reducing power:- Both alkali metal

and hydrogen act as reducing agent.



3. Reaction with oxygen:- Hydrogen burns in oxygen to form its oxide



Alkali metal also burns vigorously when heated in oxygen.



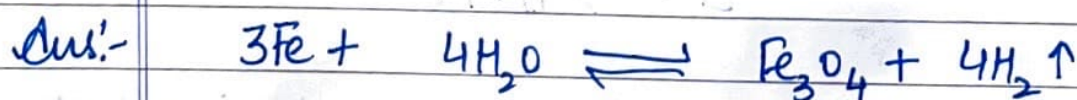
4. Oxide formation:- Sodium & hydrogen forms peroxide



Q:-4 why hydrogen is called inflammable gas?

Ans:- Because it is combustible.

Q:-5(a) Write Reaction of steam with red hot iron.



Q:-5(b) Why is this reaction considered a reversible reaction

Ans:- Because it can attain from both sides.

Q:-5(c) How can reaction proceed continuously

Ans:- By increasing temperature and concentration of reactant.

Q:-6 Why are zinc and aluminium (Al) considered to have a unique nature. Give balance equation to support your answer?

Ans:- Zn and Al reacts with steam only and become unactive after sometime due to formation of oxide layer on the surface of metal.

