

TENDER HEART HIGH SCHOOL, SEC-33B, CHD

CLASS - VIII

CHAPTER - 2

SUBJECT - CHEMISTRY

TEACHER - ANAMIKA

4.11.24

Types of Chemical Changes

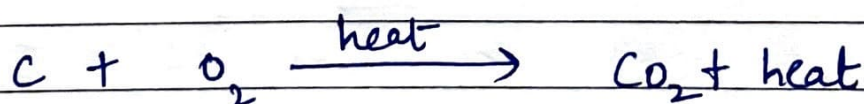
Chemical reaction involves the rearrangement of atoms among the reactant molecules to form the product molecule with different proportion.

Chemical reaction have been classified into following different types:-

1. Direct combination
2. Decomposition
3. Displacement
4. Double decomposition

1. Direct Combination:- A reaction in which two or more substances combine together to form a single substance

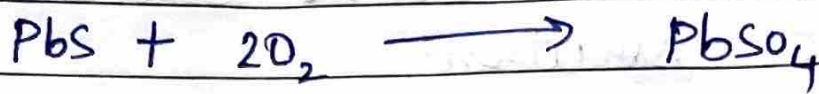
example:-



Experiment:-

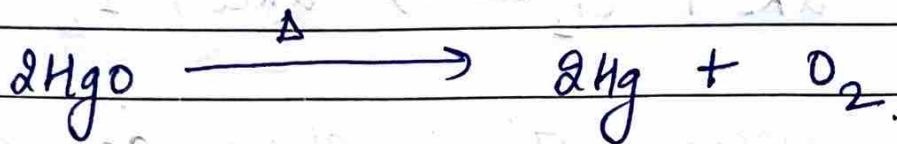
1. Take some black lead sulphide in test tube and heat it. The black lead

Sulphide react with oxygen to form white lead sulphate.



2. Decomposition:- The chemical reaction in which a compound split into two or more simpler substances is called decomposition reaction.

Example:-

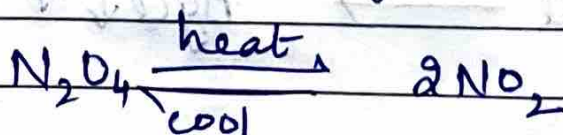


Digestion of food by our body is an example of decomposition reaction.

Fats and oils are ~~decomposition~~ to fatty acid and finally oxidized by respiration into CO_2 and H_2O .

Experiment:-

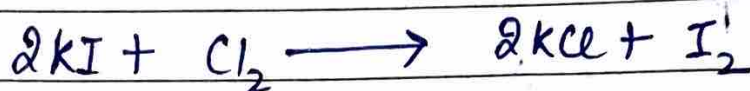
1. On heating, N_2O_4 changes to NO_2 a reddish brown gas. On cooling, NO_2 changes into original compound.



3. Displacement:- It is chemical change in which a more active element displace a less active element from its salt solution.

Experiment:-

1. Pass Cl_2 gas through a solution of KI. The colourless solution turns yellow brown as Iodine is formed.



4. Double decomposition:- chemical change in which two compounds in a solution react to form two new compounds by mutual exchange of radical.

Example:-

1. Take a solution of silver nitrate in a test tube and add dil HCl or a solⁿ of NaCl.

