TENDER HEART HIGH SCHOOL, SEC-33B, CHD.

CLASS- VIII SUBJECT- CHEMISTRY
CHAPTER- Revision TEACHER- ANAMIKA

Q:- In the formation of (i) oxygen molecule (ii) Carbon tetrachloride molecule, state the following:

a) Electronic Configuration of nearest inert gas attained
b) How many electrons are shared/ transferred in bond formation?
c) Which type of bonds these compounds form?
d) Draw their orbital diagram?

Ans:- a) (i) Oxygen molecule attain Configuration of neon.

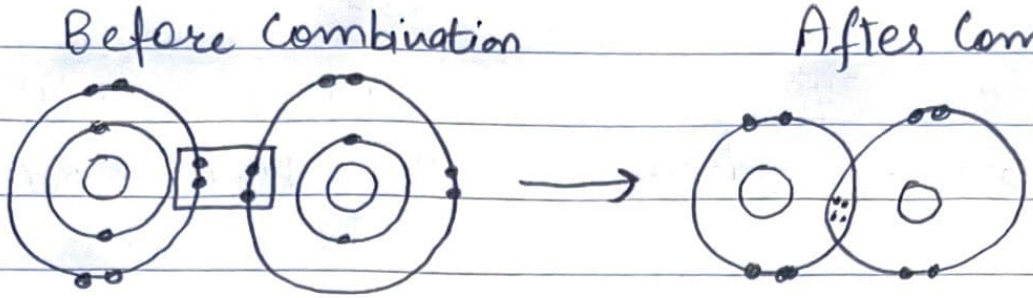
(ii) In carbon tetrachloride, carbon attain Configuration of neon and Chlorine attain Configuration of Argon.

b) (i) In oxygen molecule 2 electrons are shared by each atom

(ii) In Carbon tetrachloride, Carbon shared four electrons and chlorine shared one electron each.

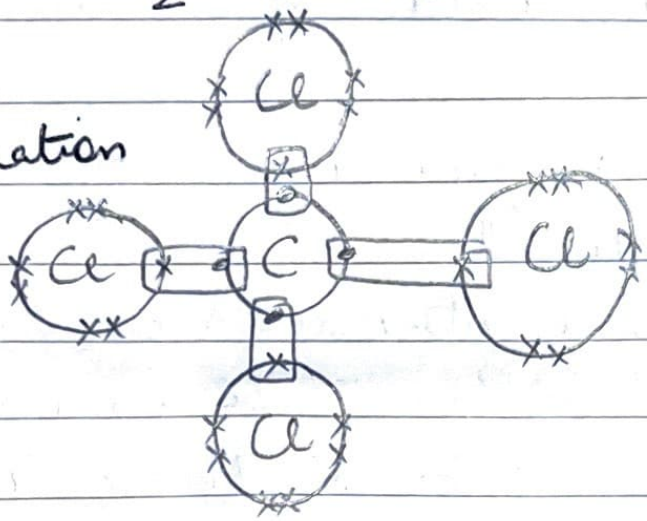
c) Covalent bond

d)

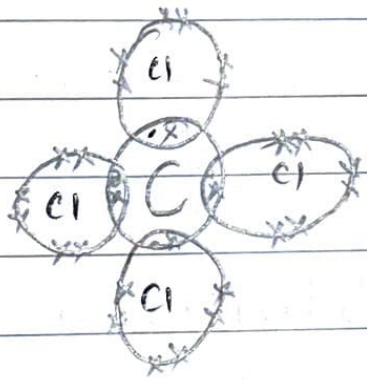


O_2 molecule

Before combination



After combination



Q:- The valency of an element A is 3 and that of B is 2. Write the formula of the compound formed by combination of A and B?

Ans:

Element A B

Valency 3 2

Interchange of Valency

A_2B_3

Q3 The formula of sulphate of an element M is $M_2(SO_4)_3$. Write the formula of its.

a) Chloride b) Oxide

c) phosphate d) Acetate

Ans:- Acc. to formula of sulphate, Valency of M is 3.

(a) Element M Cl

Valency 3 1

Interchange of Valency

MCl_3

b) Element M O

Valency 3 2

Interchange of Valency

M_2O_3

c) Element | M | PO_4
Valency | 3 | 3

Interchange of Valency

$M(PO_4)$

d) Element | M | CH_3COO
Valency | 3 | 1

Interchange of Valency

$M(CH_3COO)_3$