

TENDER HEART HIGH SCHOOL, SEC-33B, CHD.

CLASS- VIII

CHAPTER- Revision

SUBJECT- CHEMISTRY

TEACHER- ANAMIKA

Q:- An atom of an element A has 6 electrons in its M shell.

a) Write the electronic configuration of element A

Ans:-  $A \rightarrow K-L-M$   
2 8 6

b) What is the atomic number of element A?

Ans:- 16

c) Is it a metal or a Non-metal?

Ans:- Non-metal

d) What type of ion will be formed by an atom of element A? Write the symbol of ion formed?

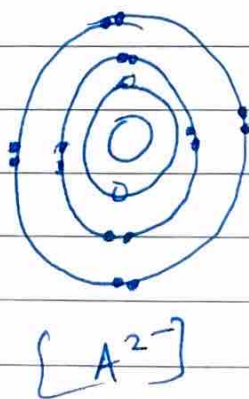
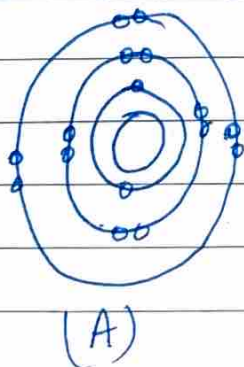
Ans:-  $A^{2-}$

e) What could element A be?

Ans:-

f) Draw the orbital structure of A and its ion

Ans:-



Q:- An element X has 2 electrons in its M shell, it forms bond with an element Y which has 7 electrons in its 3rd orbit.

a) Write the formula of the compound formed.  
Ans:-  $XY_2$

b) Which nearest inert gas electronic configuration will element X and Y acquire  
Ans:- X - acquire Neon noble gas configuration  
Y - Argon noble gas configuration

c) Show using an orbital diagram the formation of compound between X and Y.

