

TENDER HEART HIGH SCHOOL, SEC-33B, CHD

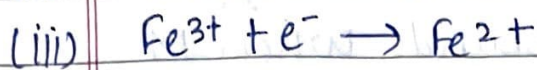
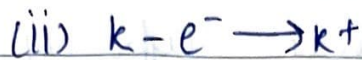
Class - VIII

Chapter - 4

Subject - Chemistry

Teacher → Anamika

Q:-1 Identify the following reaction as either oxidation or reduction?



Ans:- (i) $O + 2e^- \rightarrow O^{2-} \rightarrow$ Reduction

(ii) $K - e^- \rightarrow K^+ \rightarrow$ Oxidation

(iii) $Fe^{3+} + e^- \rightarrow Fe^{2+} \rightarrow$ Reduction

(iv) $Zn \rightarrow Zn^{2+} + 2e^- \rightarrow$ Oxidation

Q:-2 An element A has 1 electron in its first shell. It combines with element B having 7 electrons in its third shell. What type of bond is formed?

Ans:- A $\rightarrow$ K L
1

B $\rightarrow$ K L M
2 8 7

type of Bond formed between them is Polar covalent bond.

Q3:- If electrons are getting added to an element Y, then:-

a) what charge will Y have?

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b) Is Y getting oxidized or reduced?

Ans:- a) Y will take -ve charge

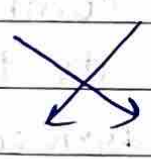
b) Y gain the electron. Therefore, It is reduced.

Q:- Metal M forms a chloride with the formula MCl_2 . What type of bond is in MCl_2 ? Write the formula of the compound when M combines with Sulphur, oxygen & nitrogen?

Ans:- M forms a chloride of formula MCl_2
 $\therefore$ M has valency 2.

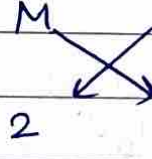
The type of bond in MCl_2 is ionic/Electrovalent bond.

Formula with Sulphur

Element M		S
valency 2		2

$\therefore$ formula is M_2S_2 or MS

Formula with oxygen $\rightarrow$

Element M		O
valency 2		2

$\therefore$ formula is M_2O_2 or MO

Formula with nitrogen Element M N

Valency

2

3

Formula $\rightarrow M_3N_2$