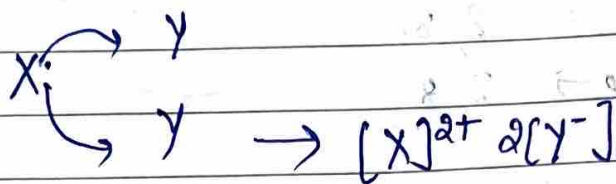
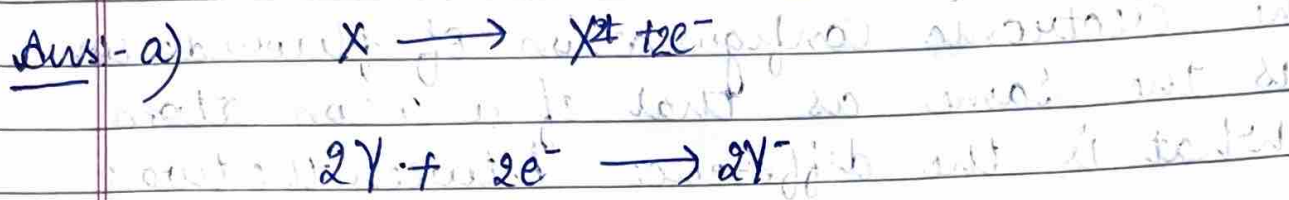
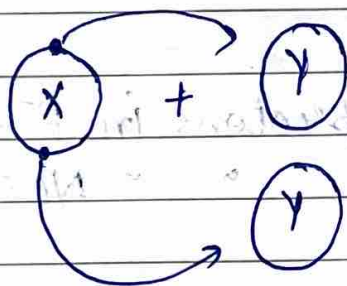


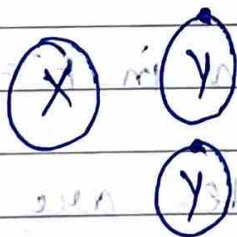
- Q:1 a) In the formation of compound XY_2 , an atom X gives one electron to each Y atom. What is the nature of bond in XY_2 ?
- b) Draw the orbit structure of this compound (XY_2)



b)



↓



- Q:2 An atom X has electronic Configuration 2, 8, 7. It combines with Y having 1 electron in its outermost shell.

- a) What type of bond will be formed between X and Y?
- b) Write the formula of the compound formed?

Ans:- a) $X \rightarrow 2, 8, 7$, $Y = 1$

X is non metal Y is metal. The bond between them is ionic/Electrovalent bond.

b) YX

Q-3 The electronic Configuration of fluoride ion is the same as that of a neon atom. What is the difference between the two?

Ans:- $F^- = 2, 8.$

$Ne \rightarrow 2, 8$

At. No. of $F^- = 9$

" " " $Ne = 10$

$\therefore$ No. of protons in $F^- = 9$

" " " $Ne = 10$

Therefore, F^- and Ne are differ in number of protons