

TENDER HEART HIGH SCHOOL, SEC-33B, CHD.

CLASS- VII
CHAPTER- 4

SUBJECT- CHEMISTRY
TEACHER- ANAMIKA

30.09.24

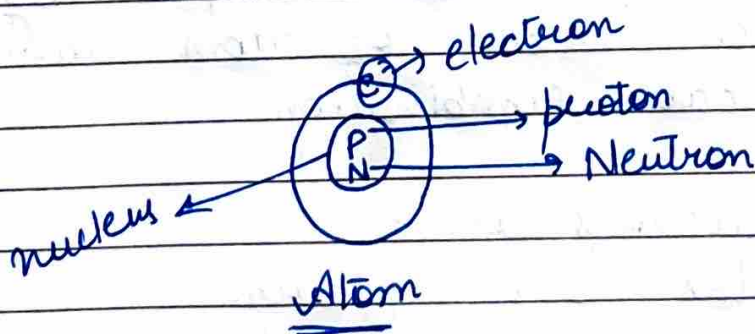
Good morning to all!

Atoms of the most elements are composed of still smaller particles known as fundamental particles or subatomic particles. They are protons, electrons & neutrons.

Protons are positively charged particles present in an atom.

Electrons are negatively charged particles present in an atom & its mass is

Neutrons are particles with no electrical charge. The central part of an atom is called nucleus, which contains both protons & neutrons.



Shells or orbit around the nucleus.

Atomic Number:- Refers to number of protons present in an atom. It is denoted by 'Z'

Mass Number:- Mass number is the sum of number of protons + neutrons present in the nucleus of an atom. It is denoted by 'A'

Mass Number (A) = Number of protons + Number of neutrons

Atomic Mass:- The mass of an atom is known as Atomic mass.

Relative atomic Mass is the mass of an element as a multiple of the standard atomic mass unit.

Valency It is the number of electron donated or accepted by valence shell of atom during chemical combination.

The Questions are:-

Define the following term:-

Atom, Molecule, Valency